relationship having parameters consistent with the postulated processes. Although details of the protonation reactions of TPAB's are not yet well understood, the present study provides a base for further exploration of the chemistry, structure, and physical properties of these interesting systems.

Experimental Section

TPAB's were synthesized according to literature procedures.[8] New TPAB's (*p*-methyl and *p*-fluoro) exhibited satisfactory elemental, ^1H NMR, and IR analyses. UV-vis spectra were recorded on a Hitachi 100-80 spectrometer equipped with a thermostated cell holder. pH measurements were made on a Corning 125 pH meter with a Metrohm 6.0216.0 combination electrode calibrated in aqueous solution using standard buffers (Fischer, Gram-Pac).

pH-Dependent UV-Visible Spectrometric Measurement. Solutions of TPAB's (ca. 1.0×10^{-5} M) were prepared by weight and sufficient $NaClO_4$ was added to attain an ionic strength of 0.50 (to minimize deviations due to ionic effects) in 90% methanol/water (100 mL of doubly distilled water to 1 L volume with distilled absolute methanol). After calibration, the pH electrode was allowed to equilibrate in 90% methanol/water for ca. 2 h prior to use. pH_{obs} values of the TPAB solutions were adjusted by using concentrated perchloric acid and/or tetraethylammonium hydroxide and UV-vis spectra were recorded at 25.0 ± 0.1 °C. pH_{obs} readings were converted to $[H^+]$ by using the relationships $[H^+] = (10^{-pH})/\gamma_{\pm}$ and $pH = pH_{obs} - \partial$, where ∂ is a correction factor for solvent effects[10,14] and the activity coefficient, $\gamma_{\pm}$, at an ionic strength of 0.5 was interpolated from the data of Akerlof.[15] Values of acid dissociation constants, K_a's, were obtained by using a nonlinear least-squares program to fit absorbance, A_j, vs $[H^+]$ data to the equation $A_j = \{A_L + (A_{HL}/K_a)[H^+]\}/(1 + [H^+]/K_a)$, where A_L and A_{HL} are the limiting absorbance values of the non- and monoprotonated forms of the TPAB, respectively.

Tosylate Adduct of Nonsubstituted TPAB (1). A solution of 54 mg (0.28) mmol) of *p*-toluenesulfonic acid hydrate in 5 mL of acetone was added to a solution of 100 mg (0.28 mmol of nonsubstituted TPAB in 5 mL of acetone. After swirling to mix, the solution was allowed to stand at room temperature overnight in a closed vial. The resulting solid that formed was collected by filtration, washed once with minimal cold acetone, and dried under reduced pressure to give 66 mg (45 %) of **1** as air stable, dark red crystals: mp 239–41 °C. Anal. Calcd for $C_{31}H_{29}N_3SO_3$: C, 71.10; H, 5.58; N, 8.02. Found: C, 71.05; H, 5.60; N, 7.96. IR (KBr): 3240 (NH), 1640 and 1595 (C=N) cm^{-1}.

Acknowledgment. We thank M. Craig, D. Guerrero, Y. Liang, and A. Schovanec for procedural assistance and help in the acquisition of spectral data. We are grateful to the National Science Foundation for support of D.A. through a Summer Undergraduate Research Fellowship.

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Additions and Corrections

Vol. 50, 1985

Steven H. Bertz. Tetramethyl 3,7-Dihydroxybicyclo[3.3.1]nona-2,6-diene-2,4,6,8-tetracarboxylate: A Useful Companion to Meerwein's Ester. Topological Analysis of Bicyclo[3.3.1]nonane Synthesis.

Page 3585. The next-to-last sentence of the first paragraph in the Results section should read "The former (see Experimental Section for spectral parameters) fits (*E*)-3-methoxy-2-propenal, whereas the latter fits the symmetrically H-bonded *Z*-isomer of **2**."

Page 3591. The second and third sentences of the first procedure in the Experimental Section should be "The ^1H NMR spectrum of a sample extracted into $CDCl_3$ had ceased to change by this time: ^1H NMR ($CDCl_3$) δ 3.82 (s, 3 H), 5.65 (dd, J = 8, 13 Hz, 1 H), 7.43 (d, J = 13 Hz, 1 H), 9.45 (d, J = 8 Hz, 1 H), which is consistent with (*E*)-3-methoxy-2-propenal.[57] Approximately 20% of another compound was also present, the ^1H NMR spectrum of which matched that of **2**.[58,59]"

See also reference 60.

Acknowledgment. I thank Prof. H. Quast (Würzburg) for bringing to my attention the need for this correction.

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(58) Bothner-By, A. A.; Harris, R. K. *J. Org. Chem.* **1965**, *30*, 254.

(59) George, W. O.; Mansell, V. G. *J. Chem. Soc. B* **1968**, 132.

(60) Bertz, S. H.; Dabbagh, G. *J. Org. Chem.* **1990**, *55*, 5162.

Vol. 54, 1989

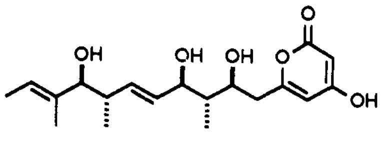
Hans P. Beutelman, Linfeng Xie, and Williams H. Saunders, Jr.*. Deuterium Isotope Effects and the Mechanism of Kinetic Enolate Formation.

Page 1704. Numbers in the last column of Table I were incorrectly transcribed and should read from top to bottom: 3.13 ± 0.29, 3.29 ± 0.31, and 6.59 ± 0.50.

Vol. 55, 1990

Hidenori Danda, Marvin M. Hansen, and Clayton H. Heathcock*. Reversal of Stereochemistry in the Aldol Reactions of a Chiral Boron Enolate.

Page 173, column 1. The structure of ACRL toxin IIIA (1) is depicted incorrectly; the correct structure is



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Apurba Datta, Hiriyakkanavar Ila,* and Hiriyakkanavar Junjappa*. Reformatsky Reaction on α-Oxo Ketene Dithioacetals: Synthesis of Substituted and Fused Ethyl 2-Hydroxy-6-(methylthio)benzoates, 6-(Methylthio)pyran-2-ones, and 6-(Methylthio)-2(1*H*)-pyridone Derivatives.

Page 5591, Scheme V. Starting materials for entries 6 and 7 should be **10a**[a] and **10c**[b], respectively. Footnotes to Scheme V are as follows: [a] $R^1 = C_6H_5$. [b] $R^1 = 4\text{-}MeOC_6H_4$.